

Chemical Reactions Chemquest 28 Answer

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Introduction to Reactions During a chemical reaction, new substances are formed. Reactants are transformed into different products. Atoms are never created or destroyed, but they are rearranged.

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In both reactions a single atom replaces another atom from a compound that it is reacting with. $\text{NaCl} + \text{Ag} \rightarrow \text{AgCl} + \text{Na}$
(note: order is not important, so you could also write $\text{Na} + \text{AgCl}$) $\text{Li} + \text{Ca}_3(\text{PO}_4)_2 \rightarrow \text{Li}_3\text{PO}_4 + \text{Ca}$ (again, order is not important) CHEMICAL REACTIONS

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ChemQuest 28 $\text{LiBr} + \text{P} \rightarrow \text{Li}_3\text{P} + \text{Br}_2$ Why is bromine written as Br_2 on the right side?

CHEMICAL REACTIONS

A chemical equation represents what happens during a reaction. The following is an example of a chemical equation:

Example Equation: $\text{Ca} + \text{HNO}_3 \rightarrow \text{Ca}(\text{NO}_3)_2 + \text{H}_2$. This equation describes the reaction of calcium (Ca) with nitric acid (HNO_3) to produce calcium nitrate ($\text{Ca}(\text{NO}_3)_2$) and hydrogen gas (H_2).

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and salt: $\text{HCl} + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{NaCl}$
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Here is what a chemical reaction looks like when it is written $\text{Cu}(\text{NO}_3)_2 + \text{Zn} \rightarrow \text{Zn}(\text{NO}_3)_2 + \text{Cu}$ There is a sentence that you can write to describe what happens above: Copper(II) nitrate reacts with zinc

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metal and yields zinc nitrate and copper metal. Take a minute to notice the relationship between the sentence and the symbols in the reaction

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you learned about single replacement
reactions in which a single atom
replaces an ion from another
reactant. Study what happens in the
following reactions. They are called

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Answer

double replacement reactions.

Fig-t - LTHS Answers

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