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Solution:  $Q = [O_2(g)]^2$   $Q = 5.0 \times 10^{-2}$   
Statement: The value of  $Q$  is greater than  $K$ , so the reaction will shift to the left, toward the reactant. Tutorial 2 Practice, page 454  
1. Given: Volume,  $V = 250 \text{ mL} = 0.25 \text{ L}$ ;  $n_{\text{initial}} I_2(g) = 0.50 \text{ mol}$ ;  $n_{\text{initial}} Br_2(g) = 0.50 \text{ mol}$ ;  $K = 1.2 \times 10^2$   
Required:  $[I_2(g)]$  equilibrium;  $[Br_2(g)]$  equilibrium  
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